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The $\text{Ba}_{10}\text{S}(\text{VO}_3\text{S})_6$ Oxysulfide: 1D-Structure and Mixed Anion Chemical Bonding

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Abstract

The new oxysulfide $\text{Ba}_{10}\text{V}_6\text{S}_7\text{O}_{18}$, that can be written as $\text{Ba}_{10}\text{S}(\text{VO}_3\text{S})_6$, was prepared by solid state reaction. It crystallizes in the non-centrosymmetric $P6_3$ space group with the unit cell parameters $a=18.3018(2)$ Å and $c=8.6525(2)$ Å, $R_1 = 3.21\%$. This original phase exhibits (VO_3S) units separated by Ba^{2+} cations, the later delimit 1D-hexagonal-like cavities filled by disordered sulfur S^{2-} anions and arranged into two kinds of sulfur-deficient 1D-channels. DFT calculations were employed to get insights into the chemical bonding and parameters ruling the structure, particularly the V-O vs. V-S bonding inside the mixed anion VO_3S tetrahedra, and the contribution of the S^{2-} of the cavities. The title compound can be decomposed with three components mainly interacting by ionic bonds as follows, $\text{Ba}_{10}\text{V}_6\text{S}_7\text{O}_{18} \rightarrow [\text{Ba}_{10}]^{20+}[\text{S}]^{2-}[(\text{VO}_3\text{S})_6]^{18-}$, this description may pave the way to the design of other phases related to this system with adjusted band gap features. In particular, the effect of the $\text{V}(\text{O,S})_4/\text{Ba}$ ratio is discussed to emphasize the presence of the $[\text{S}]^{2-}$ component, in comparison with related structures such as $\text{Ba}_6\text{V}_4\text{O}_5\text{S}_{11}$ [$\text{Ba}_6(\text{VO}_2\text{S}_2)_2(\text{VS}_3\text{O})(\text{VS}_4)$], as it contributes strongly just below the fermi level with subsequent alteration of the band gap.

Keywords: Mixed anion, oxysulfide, oxychalcogenide, heteroleptic coordination, band gap engineering, building units, DFT, COHP

Introduction

The search of new functional semiconductors can greatly benefit from the possibilities offered by mixed anion systems¹ as shown from a number of examples, in particular considering oxychalcogenides in which the different chemical bonding related to O^{2-} or Ch^{2-} allows tuning the electronic structure.²⁻⁴ For instance, $LaCuOCh$ ($Ch = S, Se$ and Te) phases are p -type wide-band-gap-semiconductors,^{5,6} in which the structure is formed by the stacking of (LaO) fluorite type layers and (CuCh) anti-fluorite layers. La_2O_2S is another wide-band-gap layered oxysulfide and efficient phosphor, with versatile emission performances of the rare-earth-doped phase.^{5,7,8} Here, the fluorite (La_2O_2) slabs alternate with sulfur layers. In the later examples, the anions segregation favors a two dimensional structuration. Although O^{2-} and S^{2-} are quite different in size and chemistry, they can lead to heteroleptic coordination around a metal cation, i.e. mixed anion building units, such as in the polar phase $CaOZnS$ with distorted ZnS_3O tetrahedron (Td). A strongly polar structure is stabilized leading to finely tunable photoluminescent and mechanoluminescent properties, after Mn doping at the Zn site.⁶ For similar reasons the polar $BaGeSe_2O$ phase, another rather scarce example, displays large non-linear-optical response compared to the oxide analogue due to cooperative arrangement of the distorted mixed anion GeO_2Se_2 building units.⁹ In all these examples the polar $MO_{4-x}Ch_x$ groups are the key structural entity for the generation of striking properties with mechanisms relying on the anionic diversity.

Other oxychalcogenides illustrate well the potentialities of band gap engineering from an heteroleptic coordination at the transition metal site, such as the $A_3V(O_{4-x}S_x)_3$ ($A = Na^+, K^+$, $x = 0, 1, 2, 3$ or 4) phases, made of Alkaline cations A^+ and $V(O, S)_4$ Td with different O/S ratios leading to colors ranging from yellow to deep red upon increased sulfur content, $VO_4 \rightarrow VO_3S \rightarrow VO_2S_2 \rightarrow VS_4$ ^{10,11} Besides alkaline elements, several of the later Td cohabit together with Ba^{2+} , so far in one remarkable compound, $Ba_6V_4O_5S_{11}$, i.e. $Ba_6(VO_2S_2)_2(VS_3O)(VS_4)$. The later exhibits three of those mixed oxygen/sulfur or sulfur Td simultaneously (VOS_3 , VO_2S_2 and VS_4). A cooperative polar arrangement of VO_3S units is also found in the $Ba_3V_2S_4O_3$ phase, a quasi-1D $S=1$ magnet and strongly correlated Mott insulator,¹² where $(VS_3)_\infty$ chains built from face-sharing $V^{3+}S_6$ octahedra may be replaced by $(CrS_3)_\infty$ chains (see reference 12b) and are separated by Ba^{2+} and VO_3S entities. Note that these phases represent the unique oxysulfide cases within the broad series of isotypic chalcogenides $A_3(M_{1-x}S_3)(XS_4)$ ($A =$ rare earth ; $M = Fe,$

Cu, ... ; X= Si, Ge)).¹³ The chains provide tunable low dimensional magnetic/electronic properties within this series.

Here, we present a new structural type in which barium is associated with VO_3S *Td* as well as isolated sulfur atoms which is rather uncommon. This orange phase exhibits an experimental bandgap of 2.25 eV, that we comment in terms of electronic structure and compare to the series $A_3(\text{VO}_3\text{S})$ ($A= \text{Na}^+, \text{K}^+$) and to the phase $\text{Ba}_6(\text{VO}_2\text{S}_2)_2(\text{VS}_3\text{O})(\text{VS}_4)$. The electronic structure in this group of material should allow identifying interesting features and key parameters for the design of new optically tuned semi-conductors.

Experimental section

Synthesis: Single crystals of $\text{Ba}_{10}\text{V}_6\text{S}_7\text{O}_{18}$ were initially found in a preparation based on the precursors $\text{BaO}/\text{MnO}/\text{V}/\text{S}$ in the ratio 3:1:1:3. The later were mixed and thoroughly grounded in an agate mortar before being pressed into pellets and heated in evacuated sealed quartz tube. The preparation was carried out in an Argon dry-glove box. The temperature program consisted in heating at a rate of 50°C up to at 800°C for 96 hours then cooling at $50^\circ\text{C}/\text{h}$ down to room temperature. The final product was a mixture of various phases including the title compound. We attempted to synthesize a single phase material from a mixture of the precursors $\text{BaS}/\text{BaO}/\text{V}_2\text{O}_5$ in the ratio 7:3:3 grounded in an agate mortar before being pressed into pellets. The preparation was carried out in an Argon dry-glove box and heated in an evacuated sealed quartz tube at various temperatures (600 to 900°C) to determine the best protocol. A rate of 50°C up to the synthesis temperature was applied and the targeted temperature maintained for 96 hours. The products were grounded and pressed into pellets for additional heating using the same conditions, but no significant improvement of the purity was observed. The same observation was made when varying the cooling protocol.

The powder X-ray diffraction pattern was collected on a Bruker D8 diffractometer equipped with a linear detector Lynxeye ($\text{Cu}_{K\alpha}$) in Bragg-Brentano geometry at room temperature.

X-ray diffraction on single crystals was performed on an APEX DUO diffractometer equipped with a bi-dimensional CCD 4K detector and a $\text{Mo } K_\alpha$ source.

For the study of $\text{Ba}_{10}\text{V}_6\text{S}_7\text{O}_{18}$, **DFT calculations** were carried out by employing the projector augmented wave (PAW)^{14,15} method encoded in the Vienna ab initio simulation package (VASP)¹⁶ and the generalized gradient approximation of Perdew, Burke and Ernzerhof (PBE)

for the exchange-correlation functionals.¹⁷ The full geometry optimizations were carried out using a plane wave energy cutoff of 550 eV and 20 *k* points in the irreducible Brillouin zone. It converged with residual Hellman-Feynman forces on the atoms smaller than 0.03 eV/Å and led to a good match with the experimental structure, i.e. within a reasonable error expected for the GGA method. The relaxed structure was used for calculations of the electronic structure. For the later, the plane wave cutoff energies of 550 eV and the threshold of self-consistent-field energy convergence of 10⁻⁶ eV were used, with 128 *k* points in the irreducible Brillouin Zone. By comparison, for Ba₆V₄S₁₁O₁₁, 550eV, 10⁻⁶ eV and 24 *k* points were used, respectively. Then, for the title phase, a COHP (Crystal Orbital Hamilton Population) analysis was carried out in the framework of the LOBSTER software¹⁸⁻²⁰.

Scanning electron microscopy (SEM) experiments and EDX analysis were carried out on a Hitachi S400N.

UV-Visible measurements: The absorbance of the sample was measured from 250 to 800 nm on a Perkin Elmer Lambda 650 device.

Results and Discussion

Crystal structure: After the selection of a suitable orange single crystal, XRD data collection was carried at room temperature using MoK_α radiation ($\lambda=0.71073$ Å). The details of data collection strategy and refinement parameters are reported table 1. The structure was solved in the P6₃ space group with the unit cell parameters $a=18.3018(6)$ Å and $c=8.6525(4)$ Å, $R=0.0321$, by Charge Flipping method, and least-squares refined using the JANA2006 Software package²¹, leading to the formula Ba₁₀V₆S₇O₁₈ in good agreement with the EDS analysis from which the average ratio 10/5.8/6.3 is found for Ba/V/S (see also S3 for a SEM image and a picture of the orange colored sample). Four twinned domains with similar weight (*ca.* 25%) related to the 6/mmm Laue class and racemic components were taken in account. The atomic positions and anisotropic displacement parameters are reported in tables 2 and 3. It shows disconnected VO₃S tetrahedrons separated by Ba²⁺ cations leading to the creation of an open framework. Here, the Ba²⁺ cations delimit two kinds of 1D channel with face-sharing Ba₆

octahedral cavities that accommodate or not additional sulfur anions S^{2-} . Besides, there are three independent sites of Vanadium atoms in VO_3S coordination with 1.631 to 1.790 Å typical V-O distances and 2.145 to 2.195 Å typical V-S distances, see table 4. The barium cations array presents a certain degree of disorder with two satellite positions for Ba4 and Ba5 with complementary occupancies, i.e. Ba4a/4b (occ. 0.914(2)/0.086(2)) and Ba5a/5b (occ. 0.83(2)/0.17(2)). Those barium atoms are among the cations surrounding sulfur atoms in the cavities. Accordingly, the S4 and S5 sulfur atoms isolated in those hexagonal-like cavities are also partially occupied with a total S occupancy restrained for electroneutrality. Then two kinds of channels coexist with partially occupied BaS_6 octahedra. In the so-labelled “channel 1”, S4 (occ. 0.852(12)) atoms form infinite chains of those octahedra with local vacancies responsible for the Ba5a/Ba5b split. Here the next anionic shell around S consist of quasi planar O_3S_3 polyhedra with compatible S--S and S--O distances, see figure 1c. In the so-labelled “channel 2”, partially occupied S5 centered (occ. 0.648(12)) and empty Ba_6 octahedra alternate leading to a strongly lacunar chain, *ca.* $1/3^{rd}$ occupied. It is clear from figure 1, that the half ordered vacancies (\square) result from the \square -- S_6 second anionic shell (involved in VO_3S units) with ideal \square --S distances between 3.1 and 3.3 Å, i.e. too short for the participation of a central S anion. On the opposite, S5 show an S-- O_6 coordination with correct distances. Thus, it turns out that the orientation of the VO_3S groups rule the occupancy of the BaS_6 cavities and chains. We note that in terms of acentricity, V_2O_3S are oriented “up” and contribute to the total dipolar momentum, while V_1O_3S and V_3O_3S almost cancel each other.

Comparison with other related structural types: The titled compound is original by the co-presence of S^{2-} isolated anions in channels and $(VO_3S)^{3-}$ mixed anion Td . To get further insights on the intrinsic VO_3S bonding scheme, it is interesting to discuss the $A_3(VO_{4-x}S_x)$ ($A=Na^+, K^+$) ionic examples. These compounds, although all consisting in simple Td separated by A^+ cations, crystallize in several different space groups depending on the A^+ cations and the O/S anions ratio. For instance, $Na_3VO_3S^{22}$ phase exhibits an orthorhombic cell in the non-centrosymmetric space group $R3c$ due to the polar cooperative alignment of the VO_3S units as found in our compound. A description of the tetrahedra packing and relative orientation in different phases is given in the corresponding literature (references 10-11, 22).

Dealing with the sulfur cavities arranged in chains found in the titled compound, similar ones are found for instance in the sulfide $Ba_9Fe_3S_{11}(S_2)_2$,²³ a high pressure polymorph of Ba_3FeS_5 . In

the later, face-sharing FeS_6 octahedra are stabilized in chains and are separated by barium atoms arranged in a manner that delimit rather similar cavities than in our phase. In particular Ba_6 octahedra sharing faces are also found and are occupied by sulfur atoms which are either isolated S^{2-} anions or involved in a disulfide pair $(\text{S}_2)^{2-}$.

Powder synthesis: Our attempts to synthesize the title phase in a pure form led to its formation but always accompanied by impurities, at least in our solid state conditions although varying the temperature and the cooling protocol. In particular, the pure phase seems to be hampered by the apparition of the stable vanadate phase $\text{Ba}_3(\text{VO}_4)_2$. Nevertheless, we have performed a multiphase Rietveld refinement using Jana2006 (ref. 21) based on the single crystal determined structure, see figure 2. We find the ratio 0.634(3)/0.366(4) for the title phase/impurity, with reasonable reliability factors ($R_{\text{obs}}= 4.98\%$, $wR_{\text{obs}}= 4.75\%$, $R_{\text{all}}= 5.09\%$, $wR_{\text{all}}=4.77\%$ and $\text{GOF}= 1.38$). If our studies show that the title phase can be formed in a large amount, the multi-phased sample makes difficult physical properties measurements requiring powder samples. However, UV-Visible optical measurements were possible as the contribution of the reported vanadate (colorless) is well established in the literature.²⁴ Figure 3 shows the optical measurement after a Kubelka-Munk²⁵ transformation, where both the band gap of our phase pointed as ~ 2.25 eV and of $\text{Ba}_3(\text{VO}_4)$ ~ 3.8 eV can be evidenced. The main challenge is to avoid the formation of $\text{Ba}_3(\text{VO}_4)_2$ using alternative synthetic methods, which is out of the scope of this work. However, DFT calculations allow having some insights on the expected properties of this phase.

Chemical bonding - Theoretical calculations:

Model used for simulation: Dealing with a slightly disordered crystal structure on some of the barium and sulfur sites with partial occupancies, we approximate an ordered model that preserves the composition $\text{Ba}_{10}\text{V}_6\text{S}_7\text{O}_{18}$ by lowering the symmetry to $P3$, see table in S1. The satellite positions Ba4a / Ba4b (Ba5a / Ba5b) with complementary occupancies were reduced to one fully occupied position by selecting the largest occupied Ba4a (Ba5a) initially. In the same way, the partially occupied S4 (occ. 0.852 in a 2b site) and S5 (occ. 0.648 in a 2a site) sulfur sites in the cavities were ordered as follows: S4 site (channel 2, in which the Ba_6 octahedra forming the chain alternate as occupied by S4 or vacant) was approximated to be fully occupied while S5 (channel 1, in which the Ba_6 octahedra forming the chain contain all S5

sites) was approximated half occupied in order to introduce vacant sites as found in channel 1. For the later, as allowed in the $P3$ subgroup with the splitting of the $2a$ $(0,0,z)$ site into two $(0,0,z)$ and $(0,0,z+1/2)$ sites, a single one $(0, 0, 0.0130)$ was filled in our model. This choice for the channels represents a necessary approximation to handle our DFT calculations on this phase without further decreasing the symmetry or supercell ordering along c . We assume that the high calculation cost of these alternative models is not relevant in this study. The full optimization led to a reasonable structure compatible with the experimental one after close inspection of the structural features and interatomic distances, see table 4. The relaxed ordered structure was used to calculate the electronic structure as detailed below. The total DOS exhibits a band gap of ~ 1.91 eV, thus a good agreement is obtained regarding the orange color of our crystals and the experimental band gap (2.25 eV) with a typical underestimation usually found from GGA calculations. The calculated band gaps (eV) for the sodium vanadate and thio-vanadates with increasing sulfur ratio were reported in the literature as follows: $\text{Na}_3(\text{VO}_4)$ (3.55) \rightarrow $\text{Na}_3\text{VO}_3\text{S}$ (3.05) \rightarrow $\text{Na}_3\text{VO}_2\text{S}_2$ (2.5) \rightarrow Na_3VOS_3 (1.93) \rightarrow $\text{Na}_3(\text{VS}_4)$ (1.67). For the potassium analogues the value decreases as follows from $\text{K}_3\text{VO}_3\text{S}$ to K_3VS_4 : 3.15 \rightarrow 2.55 \rightarrow 2.10 \rightarrow 1.80, i.e. slightly higher values than for the Na-phases. Although a different method was used, FP-LAPW calculations with the code Wien2k, the similar Generalized Gradient Approximation (GGA) should lead to the same order of underestimation of the band gap than our calculations. It places our phase close to the values for Na_3VOS_3 and far from the phases containing the same VO_3S thio-oxo-vanadate anion, highlighting the effect of isolated S^{2-} species in the cavities that decrease the band gap.^{10,11}

Density of States and COHP: The total and projected densities of states are represented in figure 3. First, focusing on the VO_3S units, the Vanadium $3d$ states are lying in the conduction band higher than 2 eV with the highest contribution between 2 eV and 4 eV approximatively. In the VB, the $3d$ states of Vanadium are hybridized with O $2p$ and S $3p$ states (see also COHP). The V $3d$ states contribution is mainly centered around -4 to -3 eV. The O $2p$ strongly contribute in the same region although found up to the very top of the VB. The S $3p$ states however, although contributing in the same region than Vanadium in the VB, are more present in the top of the VB, especially a sharp density of states peak is found around -1 eV. For comparison, it is interesting to mention our calculations for $\text{Ba}_6\text{V}_4\text{O}_5\text{S}_{11}$, i.e. $\text{Ba}_6(\text{VO}_2\text{S}_2)_2(\text{VS}_3\text{O})(\text{VS}_4)$. The DOS projected on the various Td units show their tendency to

contribute higher in the VB when increasing the sulfur content (from $\text{VO}_3\text{S} \rightarrow \text{VS}_4$), see figure 4.

In addition, the Crystal Orbital Hamiltonian Population (COHP) pictures the bonding nature and strength, see figure 5. Here, it is interesting to compare V-O and V-S bonding within the VO_3S entity of $\text{Ba}_{10}\text{V}_6\text{S}_7\text{O}_{18}$. Net bonding character is found for both, and more covalent bonding is expected for V-S regarding electronegativities: 3.44, 2.58, and 1.63, for O, S and V, respectively. Accordingly, V-S bonding states are present just below the Fermi level while bonding V-O states are predominant lower in the VB. More into details, the bonding states for V-O are found in the lower VB region with a large contribution centered on ~ -3.5 eV. Oppositely, V-S bonding is of bonding nature in all VB with higher contribution just below the Fermi level, while in the same region a small but significant amount of antibonding interactions are found for V-O (from E_F down to about -1 eV). Upper the Fermi level, both V-O and V-S interactions are found antibonding.

Now, focusing on channels and species building them, we can make the following observations: 1) the S^{2-} species inside the channels exhibit more localized states (regarding the $3p$ states at the top of the VB). 2) At the very top of the VB, they are the main contributors just below the Fermi level, while S^{2-} involved in VO_3S tetrahedron are present in a wider range and slightly below (see above). 3) In particular one sharp peak is found in the vicinity of the Fermi level ~ 0 eV (last states occupied, Valence Band Maximum-VBM) and a second ~ -0.4 eV for sulfur atoms located in channels 1 (S5) and 2 (S4), respectively. This is also well visible on the band structure in which very flat bands are found just below the Fermi level, see S1.

The strength of the bonding of this S^{2-} anions with Ba is depicted by the (integral) ICOHP absolute values, i.e. average $\text{ICOHP}_{(\text{Ba-S})} \sim -1.26$ eV, and is net bonding although rather low. Here, the channels' species are weakly interacting with the opened framework. Those S^{2-} species seems to dictate the VBM nature. In particular by further decreasing the band gap width compared to $\text{A}_3(\text{VO}_3\text{S})$ exempt of such sulfur atoms in channels. Such features are found in recent examples in the chalcogenide system $\text{Ba}/\text{M}/\text{S}$ ($\text{M} = \text{In}, \text{Ga}$), in particular in $\text{Ba}_{12}\text{In}_4\text{S}_{19}$ where both isolated S^{2-} and disulfide pairs $(\text{S}_2)^{2-}$ species are found in similar face-sharing Ba_6 octahedra forming *pseudo* chains figure 6d. The later corresponds to oligomers of seven face-shared octahedra terminated by corner-shared octahedra while our compound exhibits genuine 1D-cavities. Although the disulfide pairs determine the band gap since their contributions to the electronic structure is found both at the top of the VB and bottom of the

CB, the isolated S^{2-} are also strongly contributing at the top of the VB.²⁶ It is interesting noting that these phases were obtained using an excess of sulfur in the synthetic process. Such situation with disulfide pairs (only) determining the band gap is also well described for the phase $Ba_2F_2(S_2)$.²⁷ Although the two later phases are structurally different from the title compound, they show isolated sulfur species with determinant contributions to the electronic structure around the Fermi level as found in our compound.

Actually the formula of the title compound can be decomposed with three ionic interacting components as follows, $Ba_{10}V_6S_7O_{18} \rightarrow [Ba_{10}]^{20+}[(VO_3S)_6]^{18-}[S]^{2-}$. It is interesting to highlight the role of the A (A= alkaline, alkaline-earth) cation in the structure on the stabilization of an edifice with VO_3S alone associated to it or with additional structural elements. For both steric and charge balance reasons, Ba^{2+} (compared to Na^+ or K^+) lead to the accommodation of an excess of sulfur atoms in channels considering VO_3S tetrahedra while Ba^{2+} is combined with VO_4S_{4-x} (x = integer) units only in the case of $Ba_6V_4S_{11}O_5$ (VO_3S , VS_4 , VO_2S_2 *Td*) and $Ba_3(VO_4)_2$ (VO_4 *Td*). In more details intuitive parameters that might rule the structure (entities, isolated anions or not ...) in those systems can be discussed.

Besides the balance between anions and cations and the ratio O^{2-}/S^{2-} , the interplay between the $Ba^{2+}/(VO_xS_y)^{3-}$ ratio is of course of primary importance . For $V/Ba < 2/3$, the deficiency of negative charges leads to the accommodation of excess sulfur anions S^{2-} , which is the case of our compound with $V/Ba = 0.6$. Several compositions will be further tested in the prospection of this complex phase diagram. In contrast, In the $[(VO_xS_y)^{3-} - 3/2 Ba^{2+}]$ binary system (i.e. $V/Ba = 2/3$), only two phases are reported namely $Ba_3(VO_4)_2$ and $Ba_6(VO_2S_2)_2(VS_3O)(VS_4)$ which adopt different arrangements. Then we tried to synthesize other O/S compositions such as $Ba_3(VO_3S)_2$, but no evidence of its formation could be observed. It highlights the necessary steric agreement between sizeable $(VO_xS_y)^{3-}$ groups and large Ba^{2+} cations for stable crystal structure. However, several metastable phases requiring specific synthetic routes out of the scope of this paper may be designed.

Conclusion

The new oxysulfide $\text{Ba}_{10}\text{V}_6\text{S}_7\text{O}_{18}$ crystallizes in the non-centrosymmetric $P6_3$ space group. Its original structure exhibits mixed anion building units, i.e. VO_3S tetrahedra, which are disconnected from each other and separated by Ba^{2+} cations. In addition the Ba^{2+} cations define channels by forming 1D-chains of Ba_6 octahedra, the octahedra are all occupied by S^{2-} anions in channel 1 while vacant sites are found in channel 2. Both the presence of heteroleptic tetrahedra and S^{2-} isolated in channels are rather uncommon. A DFT study, in particular the density of states and crystal orbital Hamilton population (COHP), allowed rationalizing the band gap with respect to the different anions of the structure: S^{2-} from the VO_3S units, isolated S^{2-} and O^{2-} . We found that the isolated S^{2-} species seems to dictate the top of the valence band nature thus further decreasing the band gap width compared to the sulfur from the VO_3S entities. We show that the title compound can be described as $\text{Ba}_{10}\text{V}_6\text{S}_7\text{O}_{18} \rightarrow [\text{Ba}_{10}]^{20+}[\text{S}]^{2-}[(\text{VO}_3\text{S})_6]^{18-}$, this description should pave the way to the design of band gap tuned phases. Finally, the combination of polarity and of original electronic structure features around the band gap in our phase should be at the origin of interesting properties that will be investigated in future work.

Supporting Information: The Supporting Information contain the unit cell parameters and atomic coordinates of the optimized structure (Table S1), the bands structure (Figure S1), a MEB image and a picture of the sample (Figure S2).

Crystal structure report

CCDC 1871321 contains the supplementary crystallographic data for this paper. The data can be obtained free of charge from The Cambridge Crystallographic Data Centre via www.ccdc.cam.ac.uk/structures.”

Acknowledgments

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FIGURES AND TABLES

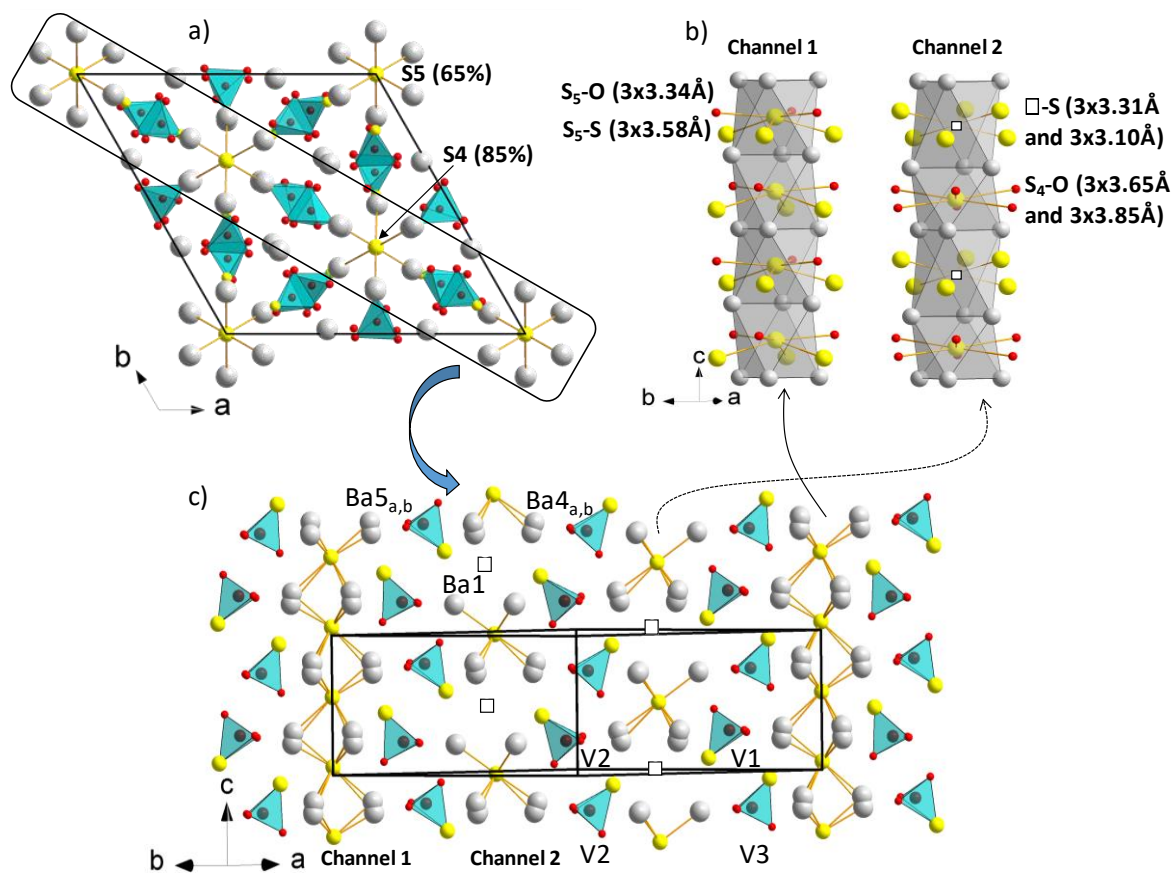


Figure 1. a) Structure of the titled compound $\text{Ba}_{10}\text{V}_6\text{S}_7\text{O}_{18}$ with Ba, S, and O represented in grey, yellow and red while the VO_3S tetrahedra are represented in blue. b) View of the channels 1 and 2 along c with a polyhedral representation, the vacancies are indicated with white squares. c) View of the structure cut shown in a).

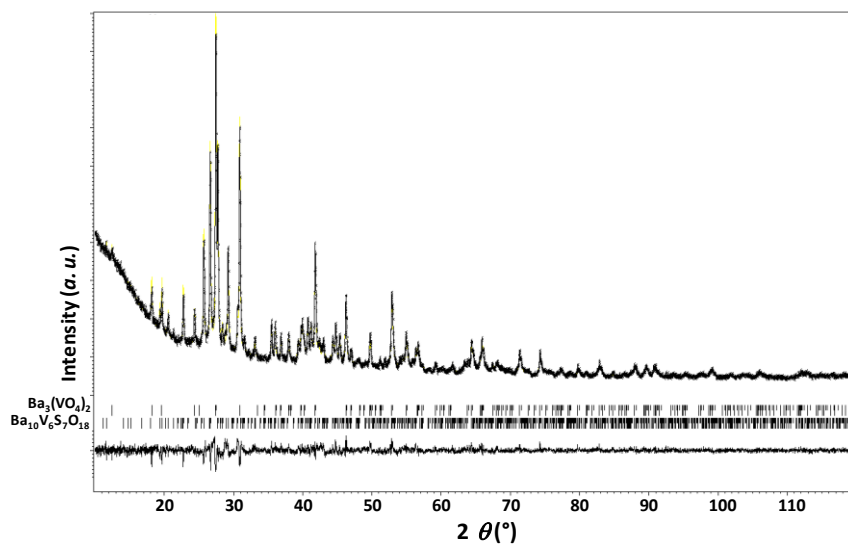


Figure 2. Multiphase Rietveld refinement of Ba₁₀V₆S₇O₁₈ with the impurity Ba₃(VO₄)₂ performed with the software Jana2006. The bottom curve is the difference between the experimental and calculated intensities. The small marks correspond to the Bragg peaks positions for Ba₁₀V₆S₇O₁₈ and the impurity Ba₃(VO₄)₂.

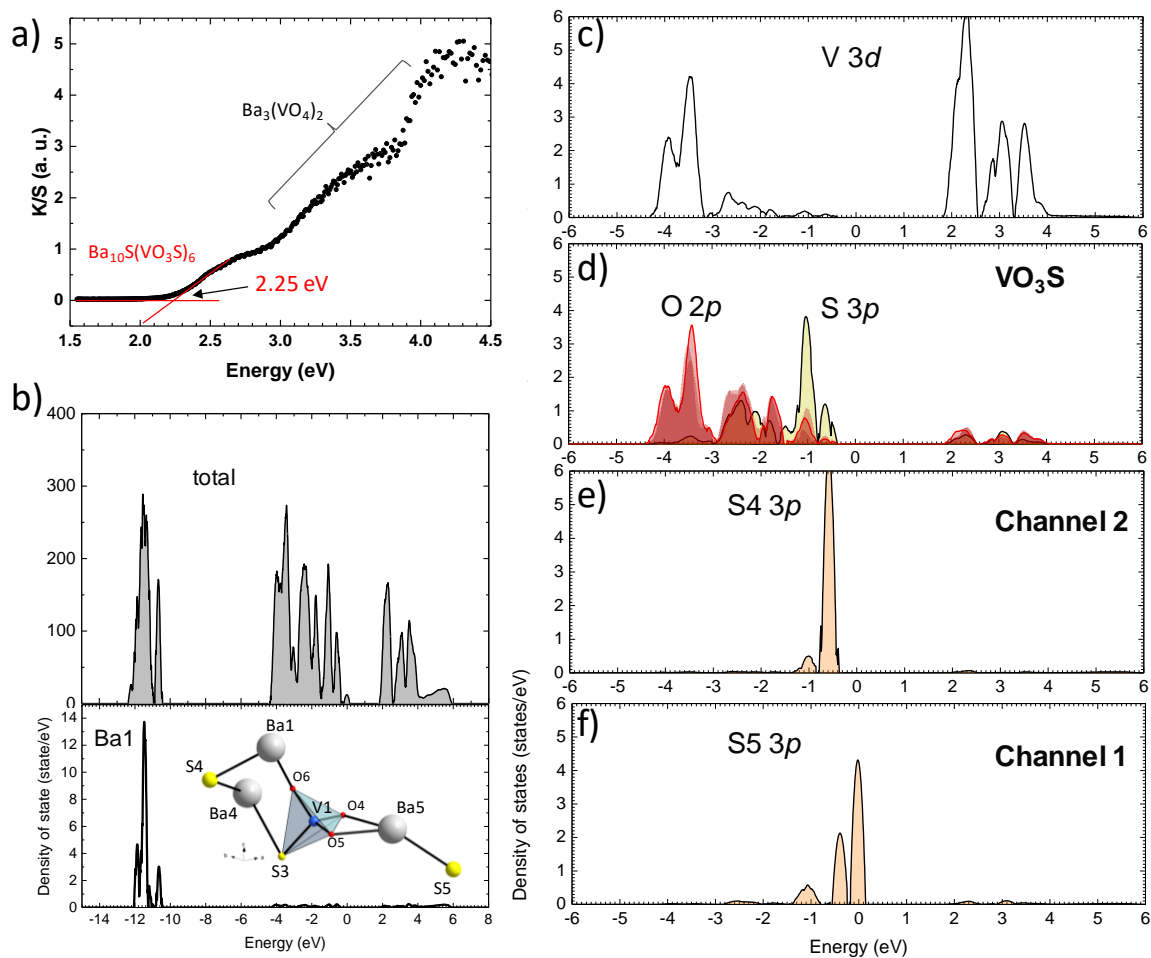


Figure 3. Optical measurements represented after a Kubelka-Munk transformation, the contribution of $\text{Ba}_3(\text{VO}_4)_2$ is shown, and the band gap of the title phase is determined as 2.25 eV. b) Total density of states and density of states projected on one Barium atom (Ba1), the other Ba atoms having similar topologies of their PDOS. A scheme of the local structure is given that includes atoms for which the PDOS is shown. c) PDOS for Vanadium and its 3d states (V1), then on d) the PDOS of its surrounding anions is given (the three oxygen atoms in red have very similar 2p states contributions while the sulfur anion 3p states in black and yellow filled curve contributes differently. S 3p states of the sulfur species within channel 2 (e) and channel 1 (f), showing their contribution just below the Fermi level and thus constituting the top of the VB. The Fermi level is set to 0.

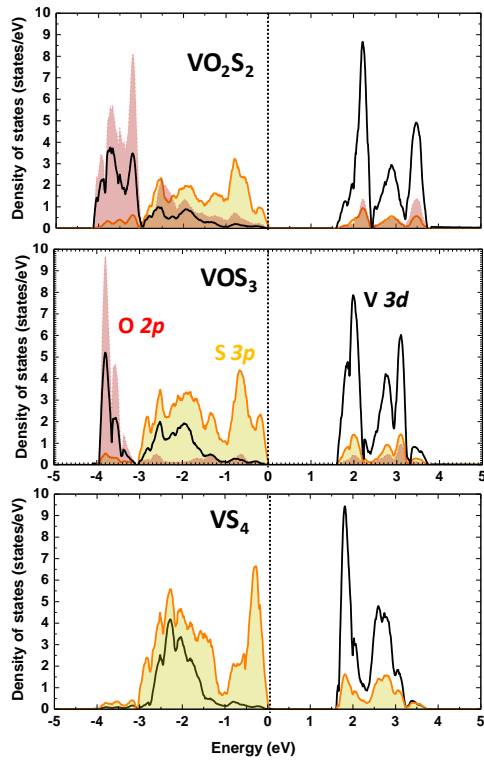


Figure 4. PDOS on the three Td entities VO_xS_y present in $\text{Ba}_6\text{V}_4\text{O}_5\text{S}_{11}$. The V 3d states are represented in black. The 2p or 3p contributions are summed per types of anions O and S, respectively, around V^{5+} to highlight the effect of the sulfur content on the VBM. The Fermi level is set to 0.

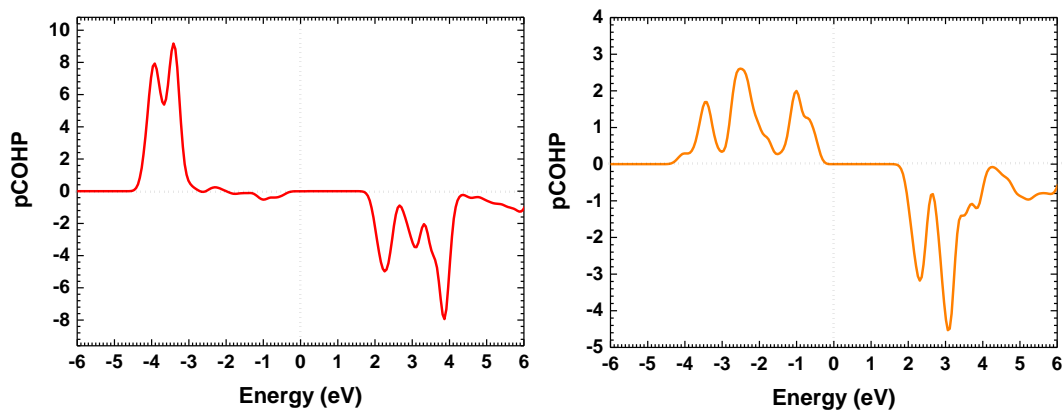


Figure 5. Crystal Orbital Hamilton populations (COHP) of the titled phase for a V-O bond (left) and a V-S bond (right) in the title phase. The Fermi level is set to 0 eV.

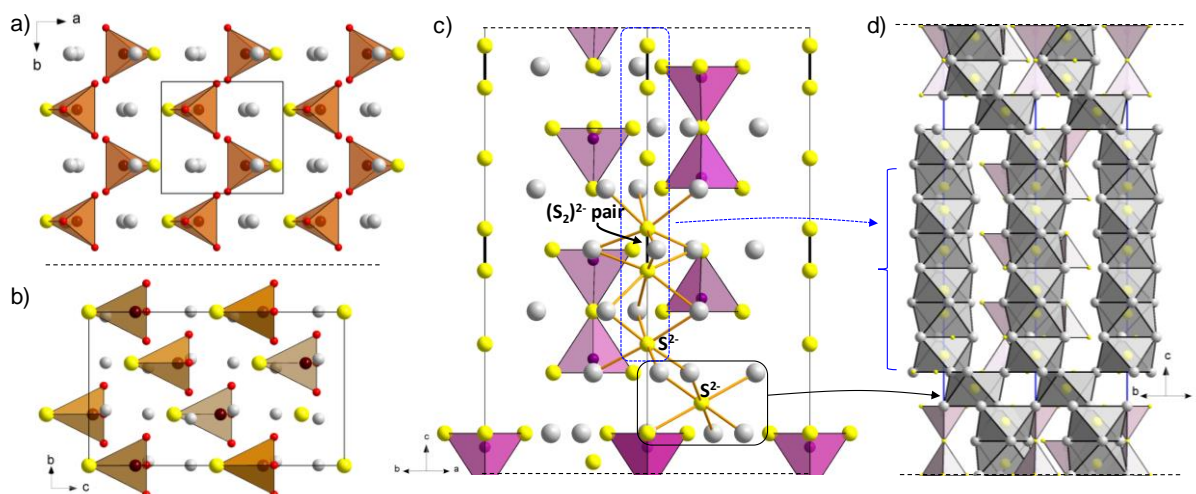


Figure 6. a) Structure of K_3VO_3S and b) Na_3VO_3S , with orange VO_3S tetrahedra and Na/K atoms in grey. c) Structure of $Ba_{12}In_4S_{19}$ with InS_4 tetrahedra in pink, the isolated sulfur S^{2-} and disulfide pairs are highlighted. d) Here we highlight the face-sharing Ba_6 octahedra (in grey) occupied by the later isolated sulfur atoms (in yellow). Broken chains (oligomers) of seven octahedra SBa_6 are formed in this phase, octahedra occupied by S^{2-} anions are alternating with two octahedra, each occupied by one of the two sulfur atoms of a disulfide pair $(S_2)^{2-}$. Focusing on the Ba/S sub-lattice, the oligomers are separated along c by a layer (in the ab plane) of SBa_6 octahedra sharing all their corners with the terminal octahedra of those oligomers.

Table 1. Single crystal XRD refinement data for the title phase $Ba_{10}V_6O_{18}S_7$

Crystal (T=298K)	
Formula	$Ba_{10}V_6O_{18}S_7$
Molar Weight ($g \cdot mol^{-1}$)	2191.4
Space group	$P6_3$
Unit cell (\AA)	$a = 18.2977(6) \text{\AA}$ $c = 8.6591(2) \text{\AA}$
Volume	2510.71\AA^3
Z	3
Data collection	
Equipment	Bruker D8 XRD900
λ (Mo $K\alpha$ (graphite monochromator); \AA)	0.71073

Density calc. (g.cm ⁻³)	4.3478
Crystal shape	platelet
Crystal dimensions (μm)	100 x 50 x 40
Color	Orange
Scan mode	ω,φ
θ (min-max) (°)	1.29 – 34.97
μ (mm ⁻¹ ; for λ Kα = 0.71073 Å)	13.655
R(int) (%)	3.73
Recording reciprocal space	-28 ≤ h ≤ 29 -29 ≤ k ≤ 29 -13 ≤ l ≤ 29
Number of measured reflections	80555
Nbr of independent reflections	5942
(I>3σ(I)), total)	5606
Refinement	
Number of refined parameters	206
Refinement method	Mean squares
Weighting scheme	Unit
R1(F) [I>3σ(I)]/R1(F ²) [all data, %]	0.0321/0.0346
wR2(F ²) [I > 3σ(I)]/wR2(F ²) [all data, %]	0.0321/0.0338
w = 1/(σ ² (F _o ²) + (0.0682P) ²)	
With P = (max(F _o ² , 0) + 2F _c ²)/3	
GOF	0.85
Max/min residual electronic density (e- /Å ³)	3.27/-1.73

Table 2. Atomic positions for Ba₁₀V₆O₁₈S₇

Atom	Wyck	S.O.F.	x	Y	z
Ba1	6c		0.24800(5)	0.50252(4)	0.1639(2)
Ba2	6c		0.01552(5)	0.33630(6)	0.4594(2)
Ba3	6c		0.32789(5)	0.34557(5)	0.4585(2)
Ba4a	6c	0.914(2)	0.58205(6)	0.16668(4)	0.2424(1)
Ba4b	6c	0.086(2)	0.5876(6)	0.1788(4)	0.1725(8)
Ba5a	6c	0.828(2)	0.07914(8)	0.16066(5)	0.7610(1)
Ba5b	6c	0.172(2)	0.0823(3)	0.1656(2)	0.6804(5)
V1	6c		0.1924(2)	0.3787(1)	0.7440(3)
V2	6c		0.0465(1)	0.5199(2)	0.1805(3)
V3	6c		0.1490(2)	0.2893(1)	0.1865(3)
S1	6c		0.1438(2)	0.5749(3)	0.3605(4)
S2	6c		0.1093(3)	0.2165(2)	0.3952(4)
S3	6c		0.2313(3)	0.4646(2)	0.5475(4)
S4	2b	0.852(12)	2/3	1/3	0.4501(6)
S5	2a	0.648(12)	0	0	0.0130(6)
O1	6c		0.1251(11)	0.2087(10)	0.0460(11)
O2	6c		0.1039(7)	0.3493(8)	0.179(2)
O3	6c		0.2511(6)	0.3549(7)	0.1805(16)
O4	6c		0.2298(7)	0.3091(6)	0.726(2)
O5	6c		0.0863(9)	0.3179(10)	0.746(3)
O6	6c		0.2493(8)	0.4407(6)	0.8958(10)
O7	6c		-0.0126(8)	0.5633(7)	0.183(2)
O8	6c		-0.0211(6)	0.4158(6)	0.208(2)
O9	6c		0.0963(6)	0.5139(7)	0.0136(12)

Table 3. Anisotropic thermal parameters (\AA^2) for $\text{Ba}_{10}\text{V}_6\text{O}_{18}\text{S}_7$

Atom	U11	U22	U33	U12	U13	U23
Ba1	0.0242(3)	0.0131(2)	0.0165(2)	0.0086(3)	-0.0028(4)	-0.0003(2)
Ba2	0.0229(3)	0.0479(4)	0.0140(3)	0.0184(3)	0.0010(3)	-0.0022(5)
Ba3	0.0294(3)	0.0375(4)	0.0139(2)	0.0220(4)	-0.0003(4)	-0.0006(5)
Ba4a	0.0244(3)	0.0116(2)	0.0155(3)	0.0064(3)	-0.0011(4)	-0.0027(2)
Ba4b	0.0244(3)	0.0116(2)	0.0155(3)	0.0064(3)	-0.0011(4)	-0.0027(2)
Ba5a	0.0242(4)	0.0138(2)	0.0360(5)	0.0073(3)	0.0043(5)	0.0007(3)
Ba5b	0.0242(4)	0.0138(2)	0.0360(5)	0.0073(3)	0.0043(5)	0.0007(3)
V1	0.035(1)	0.0124(6)	0.0138(6)	0.0082(8)	0.009(1)	0.0003(5)
V2	0.0131(6)	0.027(1)	0.0232(7)	0.0063(8)	-0.0004(6)	0.002(1)
V3	0.024(1)	0.0147(6)	0.0393(9)	0.0084(8)	-0.001(1)	-0.0028(6)
S1	0.029(1)	0.031(2)	0.026(1)	0.019(2)	-0.0090(9)	-0.006(1)
S2	0.037(2)	0.0387(1)	0.045(2)	0.022(2)	0.011(2)	0.021(1)
S3	0.031(2)	0.026(1)	0.024(1)	0.016(1)	-0.005(2)	0.0085(9)
S4	0.0093(8)	0.0093(8)	0.012(1)	0.0046(4)	0.00000	0.00000
S5	0.008(1)	0.008(1)	0.022(2)	0.0039(6)	0.00000	0.00000
O1	0.20(2)	0.13(1)	0.023(4)	0.15(1)	-0.046(7)	-0.046(6)

O2	0.011(4)	0.041(6)	0.068(9)	0.015(4)	0.002(5)	0.026(7)
O3	0.024(4)	0.046(6)	0.052(7)	0.028(5)	0.026(5)	0.044(6)
O4	0.015(4)	0.009(3)	0.09(1)	0.000(3)	-0.005(6)	-0.001(5)
O5	0.033(7)	0.052(8)	0.10(1)	0.027(6)	0.038(8)	0.037(9)
O6	0.076(8)	0.013(3)	0.007(3)	-0.005(4)	-0.002(4)	-0.002(3)
O7	0.034(6)	0.021(5)	0.09(1)	0.015(4)	-0.044(7)	-0.008(6)
O8	0.015(4)	0.012(4)	0.09(1)	0.009(3)	-0.018(5)	-0.026(5)
O9	0.025(4)	0.050(6)	0.029(4)	0.009(3)	0.010(4)	-0.003(4)

Table 4. Comparison of the refined and optimized distances (Å), the discrepancy from the experimental distances is given (%).

	Ba₅V₃O₉S_{3.5} refined	Ba₅V₃O₉S_{3.5} optimized	% discrepancy
a (Å)	18.2977	18.6550	1.95
c (Å)	8.6591	8.6994	0.46
d S4,Ba1 (Å)	3.193(3)	3.2616	2.15
d S4,Ba4 (Å)	3.195(3)	3.2541	1.84
d S5,Ba5 (Å)	2.998(4)	3.2946	9.90
d V1,O6 (Å)	1.708(9)	1.7125	0.25
d V1,O4 (Å)	1.728(2)	1.7195	0.47
d V1,O5 (Å)	1.687(1)	1.7271	2.36
d V1,S3 (Å)	2.180(4)	2.2005	0.95
d V2,O7 (Å)	1.631(2)	1.7100	4.82
d V2,O8 (Å)	1.690(9)	1.7247	2.04
d V2,O9 (Å)	1.741(2)	1.7280	0.74
d V2,S1 (Å)	2.195(4)	2.2036	0.39
d V3,O1 (Å)	1.790(2)	1.7129	4.29
d V3,O2 (Å)	1.672(2)	1.7142	2.54
d V3,O3 (Å)	1.640(9)	1.7082	4.13
d V3,S2 (Å)	2.145(4)	2.2025	2.68

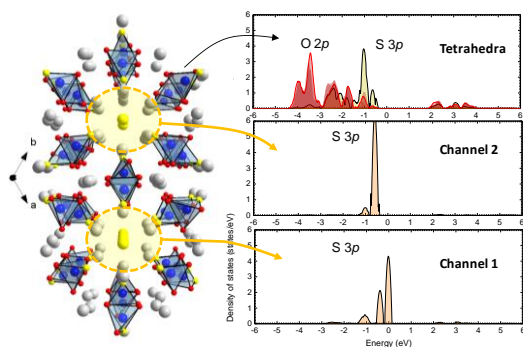
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Synopsis :

The new oxysulfide $\text{Ba}_{10}\text{V}_6\text{S}_7\text{O}_{18}$, that can be written $\text{Ba}_{10}\text{S}(\text{VO}_3\text{S})_6$, crystallizes in the non-centrosymmetric $P6_3$ space group in an original structural type with (VO_3S) mixed anion units separated by Ba^{2+} cations delimiting 2 types of deficient channels filled with isolated sulfur S^{2-} anions. DFT calculations show that the later contribute strongly just below the fermi level with subsequent alteration of the band gap. Design of systems with adjusted band gap features should arises from this work.